

Name _____

Ch 14 Study Guide CP Chemistry (Living)

$$2.1 \text{ mol OF}_2 = \text{?} \text{ molecules OF}_2$$

$$2.5 \times 10^{24} \text{ formula units Fe}_2\text{O}_3 = \text{?} \text{ mol Fe}_2\text{O}_3$$

$$0.56 \text{ mol CO}_2 = \text{?} \text{ g CO}_2$$

$$65 \text{ g CaO} = \text{?} \text{ mol CaO}$$

$$83.0 \text{ g Fe(NO}_3)_3 = \text{?} \text{ formula units Fe(NO}_3)_3$$

$$6.1 \times 10^{24} \text{ molecules P}_2\text{O}_5 = \text{?} \text{ g P}_2\text{O}_5$$

Calculate the percent of **each element** in $(\text{NH}_4)_2\text{C}_2\text{O}_4$.

Calculate the percentage of water in $\text{BaCl}_2 \cdot 9\text{H}_2\text{O}$.

How many grams of lithium can be isolated from 25.0 g of, Li_3PO_4 ?

What is the law of definite composition?

If 3.14 grams of a hydrate of LiF are heated and 2.10 g of anhydrous salt remain, what is the percentage of water in the hydrate?